Std 10 (English)   <u>Science And Technology - I</u>

Time 1HRS	Chapter 2nd chp	Marks 20

Q.1   Multiple Choice Questions   **1**

1   Mendeleev's periodic table was organized on the basis of following properties.
  a. Only physical properties
  b. Only chemical properties
  c. Both physical and chemical properties
  d. None of the above

Q.2   Find the odd one out   **1**

1   Cal/g, cal/g°C, kcal/kg.°C, erg/g.°C

Q.3   Match the pair   **1**

1

Column "A"	Column "B"
i. Elements with atomic number 19	a. Group 18
ii. Element Q in period 3 which has the highest electron affinity	b. Group 1
	c. Group 17
	d. Group 13

Q.4   State True or False   **1**

1   The number of shells in the elements of third period is three.

Q.5   Name the following   **1**

1   Give the name and symbol of elements which occupy following positions in periodic table :
Period 3Group 14

Q.6   Give scientific reasons   **4**

1   The third period contains only eight elements even though the electron capacity of the third shell is 18.

2   Metallic properties of the elements change to non metallic properties as one moves from left to right in a period of periodic table.

Q.7   Write Short Notes   **4**

1   Write short note on Mendeleev's periodic law.

2   Write short note on structure of the modern periodic table.

Q.8   Give examples   **2**

1   Give example : 3 Alkaline Earth metals and state there groups.

Q.9   Answer the following in detail (Any One)   **5**

1   Atomic number of few elements are given below 10, 20, 7, 14

  i. Identify The Elements
  ii. Identify The Group Number
  iii.Identify The Period
  iv.State Their Electronic Configuration

v. Determine The Valency of There Elements

2 An element has its electron configuration as 2,8,2. Now answer the following questions.

   i. What is the atomic number of this element?

   ii. What is the valency of this element?

   iii.What is the group of this element?

   iv.To which period does this element belong?

   v. With which of the following elements would this element resemble?(Atomic numbers are given in the brackets)

     N(7), Be(4), Ar(18), Cl(17)

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